reacted similarly to yield $62 \%$ of the boronic ester, which was hydrolyzed with sodium borate in hot aqueous ethanol to the crystalline derivative 7, $\mathrm{mp} \mathrm{104-106}{ }^{\circ} \mathrm{C}$ (from ether/petroleum ether). ${ }^{6}$ Cyclohexanone with a solution of $\mathbf{2 a}$ gave the enethiol ether 8, 83\% in crude product, $71 \%$ after treatment with aqueous ethanolic sodium borate and then sodium hydroxide to remove boron compounds and distillation. The yield of 8 from phenylthiomethyltrimethylsilane was $65 \%{ }^{11}$ Enethiol ethers were also prepared from $\mathbf{2 a}$ and butyrophenone, bp $130^{\circ} \mathrm{C}(0.1 \mathrm{~mm}), 82 \%,{ }^{6}$ and from 2 a and benzophenone, $\mathrm{mp} 69-70^{\circ} \mathrm{C},{ }^{2} 71 \%$.

One potential use of the $\alpha$-(phenylthio)alkaneboronic esters (3 and 6) is as precursors of carbonyl compounds. The pinacol boronic ester group has proved unexpectedly resistant to hydrolysis or oxidation, but cleavage of 3 a with N -chlorosuccinimide under basic conditions has given high yields of hemithioacetals or dimethyl acetals. ${ }^{12}$ Another use is the conversion of $\mathbf{1 a}$ or $3 \mathrm{a}\left(\mathrm{R}=\mathrm{C}_{4} \mathrm{H}_{9}, \mathrm{CH}_{2} \mathrm{Ph}\right)$ to $\alpha$-iodoalkaneboronic esters ( $70-77 \%$ ) with methyl iodide and sodium iodide in dimethylformamide ${ }^{3}$ for 3 days at $25^{\circ} \mathrm{C}, \alpha$-Haloalkaneboronic esters are of interest for their carbon-carbon bond-forming reactions with Grignard or lithium reagents ${ }^{13,14}$ and as precursors to boronic acids which may bind to enzymes. ${ }^{15}$ Carbon-carbon bond formation has been demonstrated with the sequence illustrated (eq 3).



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## References and Notes

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Donald S. Matteson,* Karl Arne
Department of Chemistry, Washington State University Pullman, Washington 99164
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## Additions and Corrections

Flash Photolysis of $\mathrm{Na}^{+}, \mathrm{C}^{-}\left(\mathrm{Ph}_{2} \mathrm{CH}_{2} \mathrm{CH}_{2} \mathrm{C}^{-}(\mathrm{Ph})_{2}, \mathrm{Na}^{+}\right.$. Redox Potential of 1,1-Diphenylethylene and Rate Constant of Dimerization of Its Radical Anion [J. Am. Chem. Soc., 99, 4612 (1977)]. By H. C. Wang, E. D. Lillie, S. Slomkowski, G. Levin, and M. Szwarc,* Department of Chemistry, State University of New York, College of Environmental Science and Forestry, Syracuse, New York 13210.

Because of an inadvertent error, the "wrong" Figure 7 was published. The correct figure is presented here.


Figure 7. The square root of reciprocals of the corrected slopes of the lines $1 / د$ (OD 470) vs. time (see caption to Figure 6) plotted as functions of [Trph]/[D] (triangles) or [B]/[D] (circles).

Cyclic Peptides. 17. Metal and Amino Acid Complexes of $c y$ -clo(Pro-Gly) 4 and Analogues Studies by Nuclear Magnetic Resonance and Circular Dichroism" [J. Am. Chem. Soc., 99, 4788 (1977)]. By Vincent Madison, Charles M. Deber, and Elkan R. Blout,* Department of Biological Chemistry, Harvard Medical School, Boston, Massachusetts 02115.

Page 4790, first column, 15th line from the bottom: Read "the molecular weight of cyclo(Pro-Gly),", rather than "the molecular weight of $\operatorname{cyclo}(\text { Pro-Gly })_{2}$ ".

Page 4797, first column, line 12: Change "Table IV" to read "Table II".

Page 4797, column 2, third paragraph, line 14: Change "Table VIII" to read "Table IV".

Page 4797, fourth paragraph, lines 11 and 12: Change "Figure 4 and Table X" to read "Figure 2 and Table XI".

Application of Linear Dichroism to the Analysis of Electronic Absorption Spectra of Biphenyl, Fluorene, 9,9'-Spirobifluorene, and [6.6]Vespirene. Interpretation of the Circular Dichroism Spectrum of [6.6]Vespirene [J. Am. Chem. Soc., 99, 6861 (1977)]. By Jacob Sagiv,* Amnon Yogev, and Yehuda Mazur, Department of Isotopes and Department of Organic Chemistry, The Weizmann Institute of Science, Rehovot, Israel.

Page 6868, column 1, line 1 should be: ". . . fluorene transitions lead to $z$ - and $y$-polarized exciton pairs having opposite sign in the CD spectrum, while the transversal $x$-polarized bands...".

