reacted similarly to yield 62% of the boronic ester, which was hydrolyzed with sodium borate in hot aqueous ethanol to the crystalline derivative 7, mp 104-106 °C (from ether/petroleum ether).<sup>6</sup> Cyclohexanone with a solution of 2a gave the enethiol ether 8, 83% in crude product, 71% after treatment with aqueous ethanolic sodium borate and then sodium hydroxide to remove boron compounds and distillation. The yield of 8 from phenylthiomethyltrimethylsilane was 65%.<sup>11</sup> Enethiol ethers were also prepared from 2a and butyrophenone, bp 130 °C (0.1 mm), 82%,<sup>6</sup> and from **2a** and benzophenone, mp 69-70 °C,<sup>2</sup> 71%.

One potential use of the  $\alpha$ -(phenylthio)alkaneboronic esters (3 and 6) is as precursors of carbonyl compounds. The pinacol boronic ester group has proved unexpectedly resistant to hydrolysis or oxidation, but cleavage of 3a with N-chlorosuccinimide under basic conditions has given high yields of hemithioacetals or dimethyl acetals.<sup>12</sup> Another use is the conversion of 1a or 3a ( $R = C_4H_9$ ,  $CH_2Ph$ ) to  $\alpha$ -iodoalkaneboronic esters (70-77%) with methyl iodide and sodium iodide in dimethylformamide<sup>3</sup> for 3 days at 25 °C.  $\alpha$ -Haloalkaneboronic esters are of interest for their carbon-carbon bond-forming reactions with Grignard or lithium reagents<sup>13,14</sup> and as precursors to boronic acids which may bind to enzymes.<sup>15</sup> Carbon-carbon bond formation has been demonstrated with the sequence illustrated (eq 3).

$$\begin{array}{cccc} \mathsf{PhCH}_2\mathsf{CH}-\mathsf{B}_{O}^{O} & & & \mathsf{PhMgBr} \longrightarrow \mathsf{PhCH}_2\mathsf{CH}-\mathsf{B}_{O}^{O} & & & \mathsf{H}_2\mathsf{O}_2 & & \mathsf{OH} \\ & & & \mathsf{Ph} & & & \mathsf{PhCH}_2\mathsf{CHPh} & (3) \\ & & & \mathsf{Ph} & & & (\mathsf{slow}) \\ & & & & (\mathsf{slow}) & & (52\%) \end{array}$$

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## **References and Notes**

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## Additions and Corrections

Flash Photolysis of Na<sup>+</sup>, C<sup>-</sup>(Ph)<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>C<sup>-</sup>(Ph)<sub>2</sub>,Na<sup>+</sup>. Redox Potential of 1,1-Diphenylethylene and Rate Constant of Dimerization of Its Radical Anion [J. Am. Chem. Soc., 99, 4612 (1977)]. By H. C. WANG, E. D. LILLIE, S. SLOMKOWSKI, G. LEVIN, and M. SZWARC, \* Department of Chemistry, State University of New York, College of Environmental Science and Forestry, Syracuse, New York 13210.

Because of an inadvertent error, the "wrong" Figure 7 was published. The correct figure is presented here.



Figure 7. The square root of reciprocals of the corrected slopes of the lines  $1/\Delta(OD 470)$  vs. time (see caption to Figure 6) plotted as functions of [Trph]/[D] (triangles) or [B]/[D] (circles).

Cyclic Peptides. 17. Metal and Amino Acid Complexes of cyclo(Pro-Gly)<sub>4</sub> and Analogues Studies by Nuclear Magnetic Resonance and Circular Dichroism" [J. Am. Chem. Soc., 99, 4788 (1977)]. By VINCENT MADISON, CHARLES M. DEBER, and ELKAN R. BLOUT,\* Department of Biological Chemistry, Harvard Medical School, Boston, Massachusetts 02115.

Page 4790, first column, 15th line from the bottom: Read "the molecular weight of cyclo(Pro-Gly)1", rather than "the molecular weight of cyclo(Pro-Gly)<sub>2</sub>".

Page 4797, first column, line 12: Change "Table IV" to read "Table II".

Page 4797, column 2, third paragraph, line 14: Change "Table VIII" to read "Table IV"

Page 4797, fourth paragraph, lines 11 and 12: Change "Figure 4 and Table X" to read "Figure 2 and Table XI".

Application of Linear Dichroism to the Analysis of Electronic Absorption Spectra of Biphenyl, Fluorene, 9,9'-Spirobifluorene, and [6.6]Vespirene. Interpretation of the Circular Dichroism Spectrum of [6.6]Vespirene [J. Am. Chem. Soc., 99, 6861 (1977)]. By JACOB SAGIV,\* AMNON YOGEV, and YEHUDA MAZUR, Department of Isotopes and Department of Organic Chemistry, The Weizmann Institute of Science, Rehovot, Israel.

Page 6868, column 1, line 1 should be: "... fluorene transitions lead to z- and y-polarized exciton pairs having opposite sign in the CD spectrum, while the transversal x-polarized bands . . .".